

other documents pertaining to the Office's SAWs program or policy. The Examiner is requested to make them of record in connection with this application. The disclosure of them at this time should help alleviate future questions and problems

The Changes In The Claims

Copies of this page 4 and the subsequent pages 4, 5, and 7 pertaining to changes in the claims are attached to the end of this document.

Please change the claims in this application as indicted by the following:

Cancel all of the present claims 1,3,4,6,7,8,9,10,11 and 12 in this case and replace them by the following new claims 13,14,15,16,17,18,19,20 and 21 which are based on and hopefully improved upon the claims originally in this application.

13.(new) In a method including the step of passing an electric current through an electrolyte from an inert anode to a cathode in an electro chemical cell the improvement comprising :

said electrolyte being an acidic electrolyte consisting essentially of D₂O and an acid serving as a source of hydrogen ions,

said cathode is of an electrode formed of a metal capable of taking up hydrogen ions into the physical structure of the metal, this metal being selected from the group consisting of palladium and titanium,

the temperature of the electrolyte which is between said electrodes, the acidity of the electrolyte, and the voltage and density of said current being related so that this temperature, this acidity, this voltage and density are all concurrently effective so that as said method is practiced a greater amount of heat is produced in said cell than

would be produced in said cell if the D₂O in said electrolyte was replaced by the same quantity of "regular" water.

14.(new) The method set forth in claim 13 including:
the step of collecting gases given off by reactions involving the electrolyte during the practice of said method is carried out in an enclosed space above this electrolyte;
reacting the collected gases; and
returning the reaction products of the collected gases to said electrolyte.

15.(new) The method set forth in claim 14 wherein:
the collected gases are catalytic ally reacted above the electrolyte in said cell and are returned to said electrolyte by the action of gravity.

16.(new) The method set forth in claim 15 wherein
said acid is sulfuric acid.

17. (new) The method claimed in claim 16 wherein;
said electrolyte contains 15% by volume sulfuric acid having a specific gravity of 1.84.

18.(new) The method claimed in claim 13 wherein:
said metal is palladium.

19. (new) The method claims in claim 13 wherein:

said metal is titanium.

19.(new) The method set forth in claim 13 wherein
the step of collecting gases given off by reactions involving the electrolyte
during the practice of said method is carried out in an enclosed space above this
electrolyte;

reacting the collected gases; and

returning the reaction products of the collected gases to said electrolyte.

the collected gases are catalytic ally reacted above the electrolyte in said cell
and are returned to said electrolyte by the action of gravity.

said acid is sulfuric acid.

said electrolyte contains 15% by volume sulfuric acid having a specific gravity of
1.84.

20. (new) A method of operating an electrochemical cell including an anode
during the operation of the cell and a cathode electrically connected through an
electrolyte in which:

said electrolyte consists essentially of a mixture of D₂O and an effective
amount of an acid to serve as a source of hydrogen ions,

said cathode being formed of a metal which is capable of taking up these
hydrogen ions into the physical structure of the metal during the operation of the cell
which is selected from the group consisting of palladium and titanium,

the temperature and acidity of the electrolyte between the electrodes and the
density and current between the electrodes being related so that said electrolyte is
both heated and caused to become radioactive during the operation of said cell.